Introduction: Learning molecular energy functions

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A typical molecular mechanics force field



Example: Bond length stretching

 A bonded pair of atoms is effectively connected by a spring with some preferred (natural) length.
 Stretching or compressing it requires energy.



Example: Electrostatics interaction



- Like charges repel.
 Opposite charges attract.
- Each atom carries some "partial charge" (may be a fraction of an elementary charge), which depends on which atoms it's connected to

$$U(r) = \frac{q_i q_j}{r}$$

where q_i and q_j are partial charges on atoms i and j

Could we learn an energy function (force field)

- What if instead of writing the force field as a sum of terms each of which makes physical sensewe represent it as a large neural network?
 - We can then train that network on the results of many quantum chemistry computations
- Researchers have been working on this for over a decade, but it's picked up steam in the last couple year
- One of Tuesday's papers (Smith et al.) reports a substantial step in this direction

Two related ideas

- Another paper (Faber et al.) explores prediction of chemical properties of small molecules by machine learning
 - Learning is again based on quantum chemistry results, but there's no force field involved
- A third paper (Park et al.) discusses improvement of the Rosetta all-atom force field by fitting to a wider variety of data types

Background material

- Introduction to energy functions (force fields) from CS/CME/BioE/Biophys/BMI 279:
 - <u>http://web.stanford.edu/class/cs279/lectures/</u> <u>lecture3.pdf</u>
- Discussion of the Rosetta force fields from CS/ CME/BioE/Biophys/BMI 279:
 - <u>http://web.stanford.edu/class/cs279/lectures/</u> <u>lecture5.pdf</u>